

Assessment 06



Question 1

In the electron configuration: $1s^2, 2s^2, 2p^6, 3s^2$, which ones are the valence electrons?

- A. $1s^2$
- B. $2s^2$
- C. $2p^6$
- D. $3s^2$

Question 2

In an electron configuration, which level and sub level would follow a 3p?

A. 3s

B. 4s

C. 3d

D. 4p

Question 3

The electrons in an atom that are involved in chemical reactions are the _____

- A. core electrons
- B. valence electrons
- C. electrons with positive spins
- D. electrons that exist as pairs.

Question 4

The electron configuration $1s^2, 2s^2, 2p^6, 3s^2$ is for the element _____.

- A. magnesium
- B. boron
- C. sulfur
- D. nitrogen

Question 5

Which element has the configuration $1s^2, 2s^2, 2p^6, 3s^1$?

- A. sodium
- B. magnesium
- C. potassium
- D. phosphorus

Question 6

As you move from left to right within a period in the periodic table, the number of

- A. electrons increases
- B. electrons decreases
- C. energy levels increases
- D. energy levels decreases

Question 7

What does an orbital diagram show that an electron configuration does not show?

- A. number of electrons
- B. quantum numbers
- C. electron spins
- D. energy sub levels

Question 8

The element with one more electron than neon is potassium.

- A. True
- B. False

Question 9

Moving from top to bottom within a group or family, the number of energy levels increases.

A. True

B. False

Question 10

All of the elements in group II-A have 2 valence electrons.

A. True

B. False