

Assessment07



Question 1

Which of the following elements has the largest atomic radius?

- A. Sodium (Na)
- B. Aluminum (Al)
- C. Phosphorous (P)
- D. Sulfur (S)

Question 2

Calcium (Ca) has a larger atomic radius than bromine (Br) because calcium has _____

- A. more electrons
- B. fewer electrons
- C. more energy levels
- D. fewer energy levels

Question 3

Fluorine (F) atoms are smaller than iodine (I) atoms because fluorine has _____

- A. fewer energy levels
- B. more energy levels
- C. more electrons
- D. a higher nuclear charge

Question 4

Which element has the highest ionization energy?

- A. Fluorine (F)
- B. Chlorine (Cl)
- C. Bromine (Br)
- D. Iodine (I)

Question 5

Which element has the highest ionization energy?

- A. Carbon (C)
- B. Nitrogen (N)
- C. Oxygen (O)
- D. Fluorine (F)

Question 6

Which element has the largest atomic radius?

- A. Lithium (Li)
- B. Sodium (Na)
- C. Potassium (K)
- D. Rubidium (Rb)

Question 7

Elements with a zero electron affinity are most likely _____.

- A. alkali metals
- B. alkaline earth metals
- C. Halogens
- D. inert gases

Question 8

Which of the following atoms has the greatest metallic character?

A) Na

B) Si

C) Mg

D) Al

Question 9

Group 5A elements usually form ____ ions.

A. 3-

B. 3+

C. 2-

D. 2+

Question 10

As atomic radii increase in a group, electron affinities decrease.

A. True

B. False