

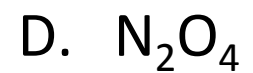
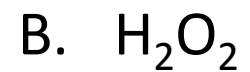
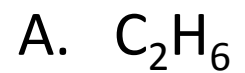
Assessment

Chemistry: Lesson 08



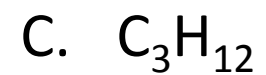
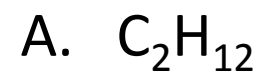
Question 1

Which of these is an “empirical formula”?



Question 2

The empirical formula for C_6H_{24} is _____.



Question 3

Molecular shapes are shown with _____ formulas.

- A. structural
- B. molecular
- C. empirical
- D. chemical

Question 4

All of the following elements are diatomic except _____.

- A. Hydrogen
- B. Nitrogen
- C. Sulfur
- D. Chlorine

Question 5

Four of the diatomic elements are in the family known as the _____.

- A. Alkali metals
- B. Alkaline earth metals
- C. Nobel gasses
- D. Halogens

Question 6

In ionic bonds, electrons are _____.

- A. transferred to the metal
- B. transferred to the nonmetal
- C. shared between two nonmetals
- D. shared between two metals

Question 7

A single covalent bond equals _____ shared electrons.

- A. 1
- B. 2
- C. 3
- D. 4

Question 8

Sodium (Na) forms covalent bonds.

A. True

B. False

Question 9

Elements and compounds are classified as:

- A. mixtures
- B. ionic substances
- C. pure substances
- D. molecular substances

Question 10

Which pair of elements should form an ionic compound?

A. Mg and Ca

B. K and S

C. N and O

D. P and Cl