

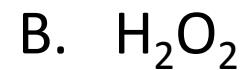
# Assessment

## Chemistry: Lesson 08



## Question 1

Which of these is an “empirical formula”?



## Question 2

The empirical formula for  $C_6H_{24}$  is \_\_\_\_\_.

- A.  $C_2H_{12}$
- B.  $CH_{18}$
- C.  $C_3H_{12}$
- D.  $CH_4$

### Question 3

Molecular shapes are shown with \_\_\_\_\_ formulas.

- A. structural
- B. molecular
- C. empirical
- D. chemical

## Question 4

All of the following elements are diatomic except \_\_\_\_\_.

- A. Hydrogen
- B. Nitrogen
- C. Sulfur
- D. Chlorine

## Question 5

Four of the diatomic elements are in the family known as the \_\_\_\_\_.

- A. Alkali metals
- B. Alkaline earth metals
- C. Nobel gasses
- D. Halogens

## Question 6

In ionic bonds, electrons are \_\_\_\_\_.

- A. transferred to the metal
- B. transferred to the nonmetal
- C. shared between two nonmetals
- D. shared between two metals

## Question 7

A single covalent bond equals \_\_\_\_\_ shared electrons.

- A. 1
- B. 2
- C. 3
- D. 4

## Question 8

Sodium (Na) forms covalent bonds.

- A. True
- B. False

## Question 9

Elements and compounds are classified as:

- A. mixtures
- B. ionic substances
- C. pure substances
- D. molecular substances

## Question 10

Which pair of elements should form an ionic compound?

- A. Mg and Ca
- B. K and S
- C. N and O
- D. P and Cl