Assessment

Chemistry: Lesson 12

Question 1

How many bonding pairs of electrons are in one molecule of ammonia (NH_3)?

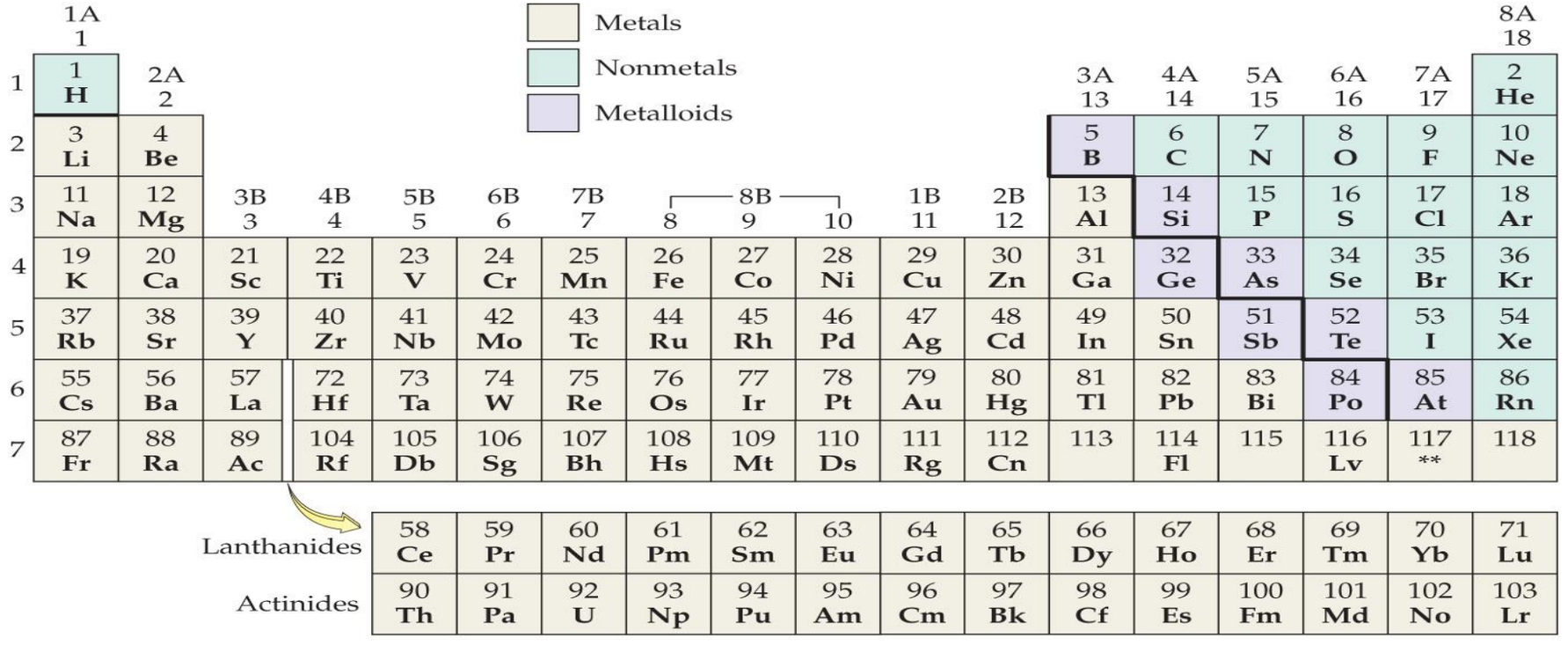
A) 2

B) 3

C) 1

D) 5

E) 0



Question 2

How many bonding pairs of electrons are in one molecule of water (H_2O)?

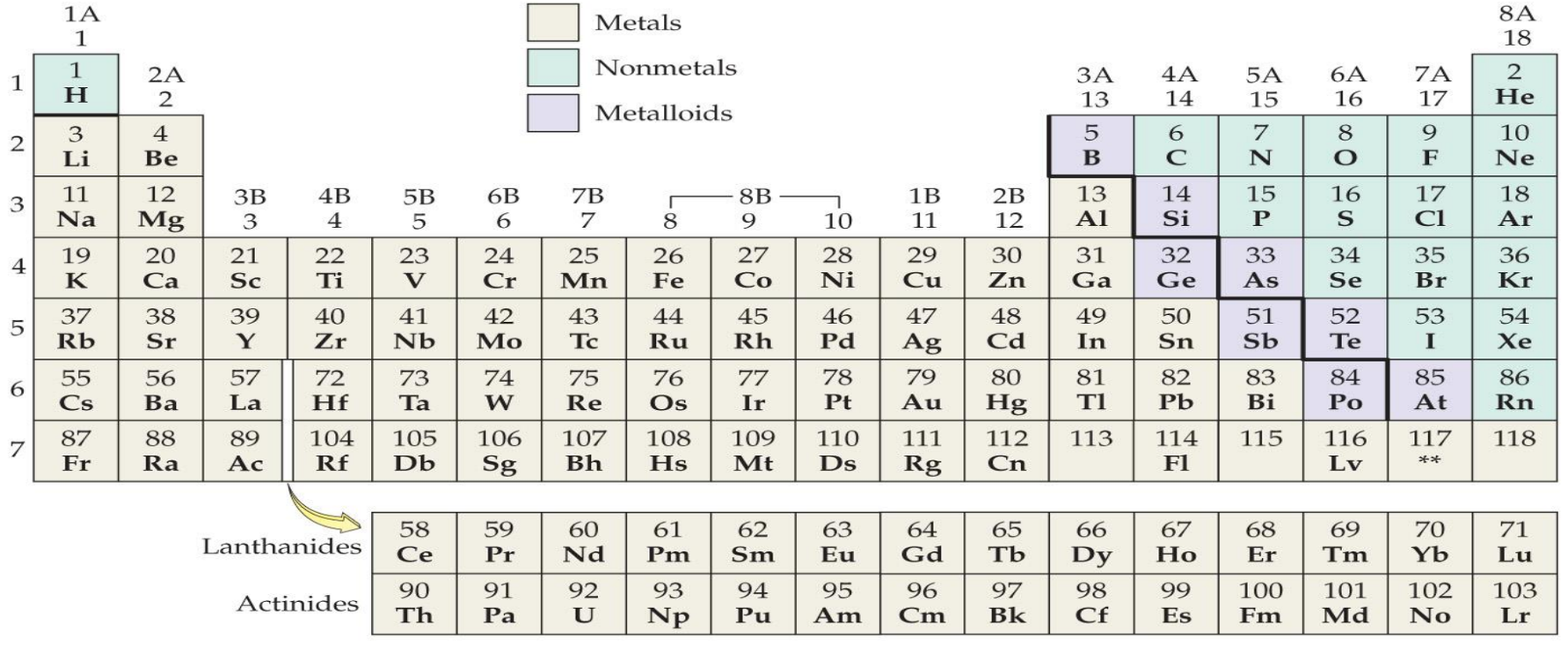
A) 0

B) 1

C) 2

D) 4

E) 6



Question 3

How many lone pairs are around the central atom in the ammonium ion?

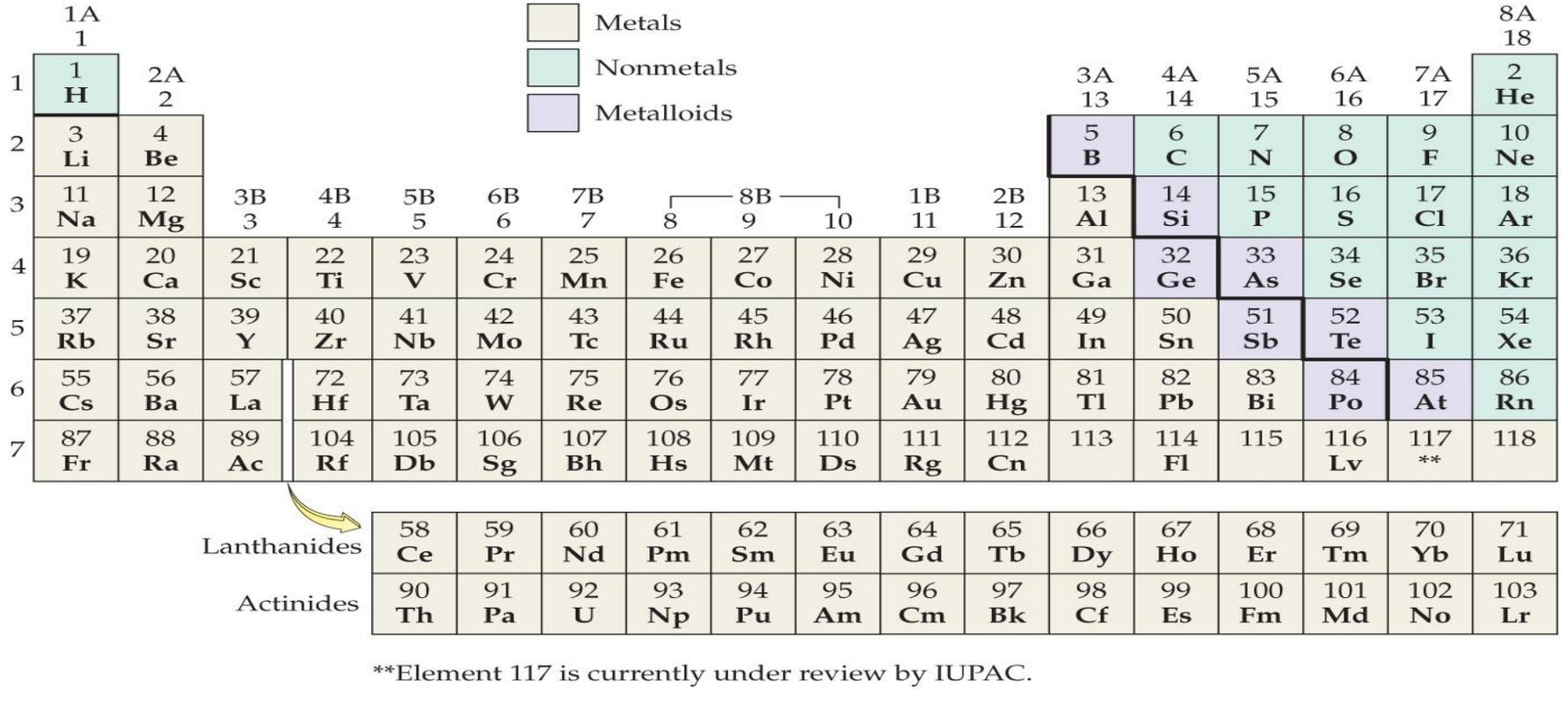
A) 1

B) 4

C) 0

D) 16

E) 12



Question 4

How many lone pairs of electrons are in sulfur atom in SO_2?

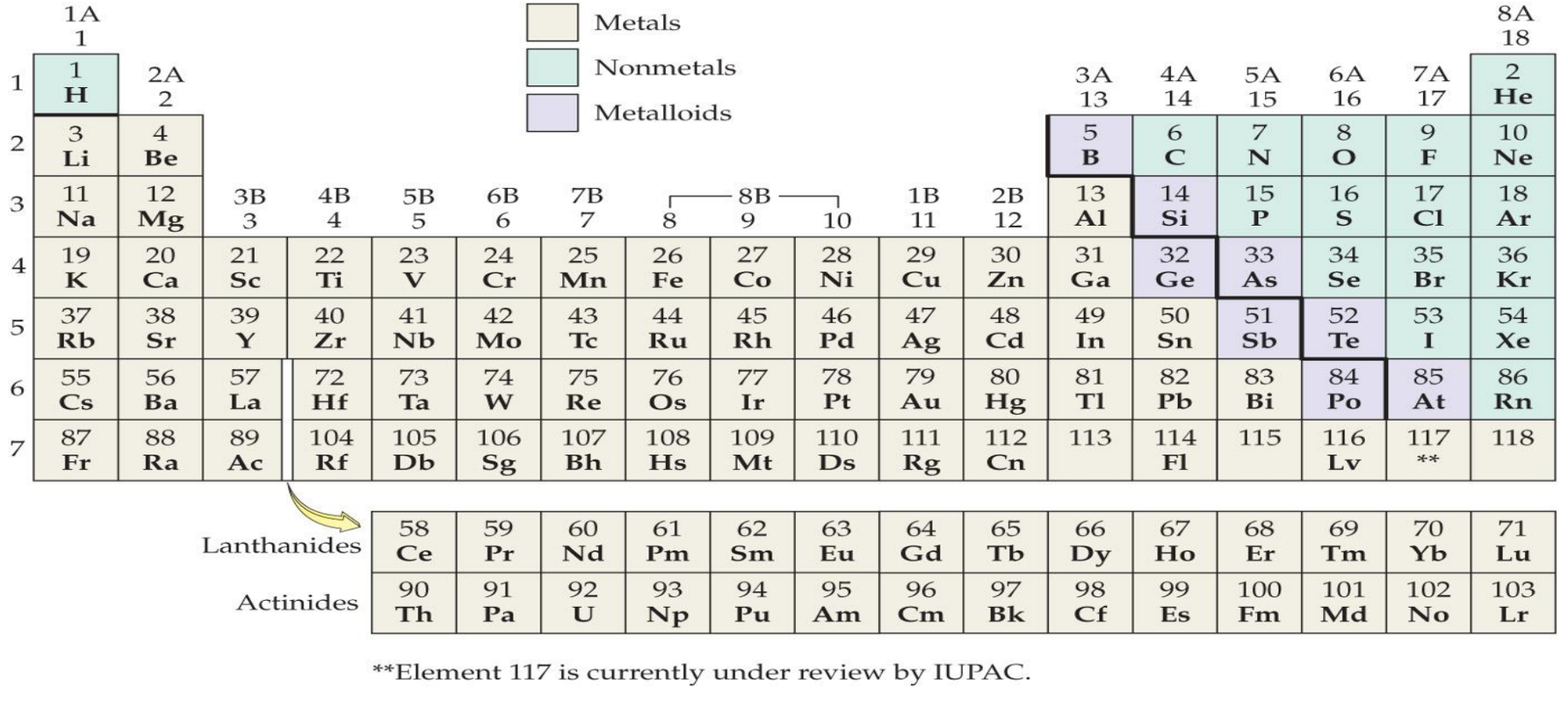
A) 2

B) 3

C) 1

D) 6

E) 0



**Element 117 is currently under review by IUPAC.

Question 5

The patterns for electronegativity in the periodic table are the same as the patterns for ionization energy.

A. True

B. False

Question 6

The most electronegative element is fluorine.

A. True

B. False

Question 7

Which is the strongest bond?

A. C – H

B. C – C

C. C = C

D. C ≡ C

Question 8

Long bonds are usually ____.

A. Strong

B. Weak

C. Triple

D. Stable

Question 9

Which should be the shortest bond?

A. C – C

B. C ≡ C

C. O = O

D. C = C

Question 10

Triple bonds tend to be _____.

A. short and weak

B. long and weak

C. long and strong

D. short and strong